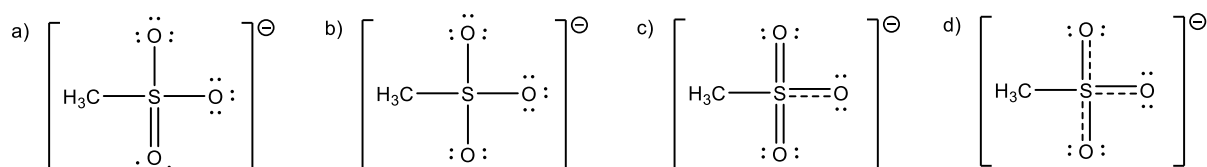
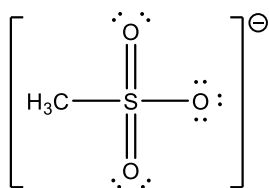


Choose the correct resonance hybrid for the following compound.



Hybrid resonance is a phenomenon that occurs when two or more resonance structures of a molecule can be described by combining the characteristics of the individual structures. In other words, hybrid resonance occurs when the electrons in a molecule are delocalized over multiple atoms, leading to a more stable molecule.

Hybrid resonance is often seen in molecules that have a resonance structure or structures, which are molecules that have more than one possible arrangement of electrons. In these cases, the actual electronic structure of the molecule is a hybrid of the different resonance structures, and the resulting electronic structure is more stable than any individual resonance structure.

The correct structure of the hybrid of the resonance is:

