



**Is  $\text{NH}_4\text{Br}$  soluble or insoluble?**

$\text{NH}_4\text{Br}$  is generally soluble in water. It is a salt that is composed of the positively charged ammonium ion ( $\text{NH}_4^+$ ) and the negatively charged bromide ion ( $\text{Br}^-$ ).

Salts of ammonium are all soluble in water.